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Synthesis, structure refinement and single-crystal elasticity of Al-bearing superhydrous phase B

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Abstract

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Dense hydrous magnesium silicates (DHMSs) with large water content and wide stability fields are a potential H_2O reservoir in the deep Earth. Al-bearing superhydrous phase B (shy-B) with a wider stability field than the Al-free counterpart can play an important role in understanding H_2O transport in the Earth's transition zone and topmost lower mantle. In this study, a nominally Al-free and two different Al-bearing shy-B with 0.47(2) and 1.35(4) Al atoms per formula unit (pfu), were synthesized using a rotating multi-anvil press. The single-crystal structures were investigated by X-ray diffraction (XRD) complemented by Raman spectroscopy, and Fourier-transform infrared spectroscopy (FTIR). Single-crystal XRD shows that the cell parameters decrease with increasing Al-content. By combining X-ray diffraction and spectroscopy results, we conclude that the Al-poor shy-B crystallizes in the *Pnn*2 space group with hydrogen in two different general positions. Based on the results of the single crystal X-ray diffraction refinements combined with FTIR spectroscopy, three substitutions mechanisms are proposed: $2\ Al^{3+} = Mg^{2+} + Si^{4+}$; $\square^{Mg2+} = 2H^{+}$; $Si^{4+} = Al^{3+} + H^{+}$. Thus, in addition to the two general H positions, hydrogen is incorporated into the hydrous mineral via point defects. The elastic stiffness coefficients were measured for the Al-shy-B with 1.35 pfu Al by Brillouin scattering (BS). Al-bearing shy-B shows lower C_{11}, higher C_{22} and similar C_{33} when compared to Al-free shy-B. The elastic anisotropy of Al-bearing shy-B is also higher than that of the Al-free composition. Such different elastic properties are due to the effect of lattice contraction as a whole and the specific chemical substitution mechanism that affect bonds strength. Al-bearing shy-B with lower velocity, higher anisotropy and wider thermodynamic stability can help to understand the low velocity zone and high anisotropy region in the subducted slab located in Tonga.

Keywords: Al-bearing superhydrous phase B, crystal structure, elasticity, Brillouin scattering, X-ray diffraction

1. Introduction

Several studies show the presence of hydrous minerals such as hydrous ringwoodite, brucite and ice-VII, as inclusion in diamonds from the deep earth interior (Pearson et al. 2014; Palot et al. 2016; Tschauner et al. 2018). This proves that the Earth's mantle is, at least locally, hydrated. H_2O can be introduced into the deep Earth by subducted slabs via hydrous minerals (Schmidt and Ulmer 2004; Ohtani 2005) or originally stored in the deep Earth from the protosolar nebula material (Hallis et al. 2015; Peslier 2020). H_2O could be incorporated in some nominally anhydrous minerals such as olivine, wadsleyite and ringwoodite. However, the H_2O-storage capacity of these minerals is limited and lower than ~3 wt.% (e.g. Mao and Li 2016). Dense hydrous magnesium silicates (DHMSs), such as Phase A, Phase E, superhydrous phase B (shy-B), phase D and phase H with large H_2O content and wide phase stability are thus potential H_2O carriers in cold or warm subducted slabs (Ohtani 2005; Nishi et al. 2014).

Mg-endmember shy-B ($Mg_{10}Si_3H_4O_{18}$) with 5.8 wt.% water, has been proposed to form in the hydrated pyrolite system at the Earth's transition zone and topmost lower mantle at pressure and temperature (P - T) conditions of subducted slabs (Pacalo and Parise 1992; Frost 1999). The Mg-endmember of shy-B coexists with ringwoodite at transition zone in the hydrous peridotite layer in subducted slab (e.g., Schmidt and Ulmer 2004; Ohtani 2005). At the depth of topmost lower mantle, shy-B will coexist with bridgmanite, phase D, $CaSiO_3$-perovskite to 30 GPa and then decompose to MgO and bridgmanite (e,g., Schmidt and Ulmer 2004; Ohtani 2005). Al_2O_3, as an important

component in the peridotite and basalt layer of subducted slab, could be incorporated into the structure of DHMSs and bridgmanite (Pamato et al. 2015; Liu et al. 2016, 2017; Kakizawa et al. 2018). Phase D and shy-B could incorporate 50 wt.% and 31.9 wt.% Al_2O_3 in their structures (Boffa Ballaran et al. 2010; Pamato et al. 2015; Kakizawa et al. 2018). The solubility of Al_2O_3 in bridgmanite is up to 21.8 mol% at 27 GPa and 2500 K (Liu et al. 2017). The partition coefficient of Al between shy-B and bridgmanite is considered to be 1.5-2.8 and increases with pressures and temperatures (Litasov and Ohtani 2003; Kakizawa et al. 2018). However, the partitioning of Al among coexisting shy-B, phase D and bridgmanite is still experimentally unconstrained.

The Mg-Si end-member of shy-B can be synthesized at high P-T in the range of 16-22 GPa and 1000-1550 ˚C (e.g. Pacalo and Parise 1992; Hazen et al. 1997; Koch-Müller et al. 2005). The structure of shy-B was first determined to be orthorhombic (space group *Pnnm*) by Pacalo and Parise (1992). The crystal structure consists of Si-O tetrahedra and octahedra with 2:1 ratio as well as Mg-O octahedra (Pacalo and Parise 1992). Koch-Müller et al. (2005) proposed that shy-B displays two different polymorphic modifications, a high-temperature form with space group *Pnnm* and a low-temperature form with space group *Pnn*2. Kakizawa et al. (2021) confirmed these findings. Different experimental techniques, including FTIR spectroscopy, nuclear magnetic resonance (NMR), single-crystal X-ray diffraction (XRD) and neutron diffraction all reveal the presence of 2 distinct hydrogen positions in the structure of shy-B which makes the structure to be compatible with *Pnn*2 space group (Koch-Müller et al. 2005; Xue et al. 2008; Trots et al. 2013; Koch-Müller et al. 2014).

The incorporation of F^-, Ti^{4+}, Fe^{2+} and Al^{3+} in the structure of shy-B affects the physical properties of this phase. The incorporation of F^- could increase the

incompressibility, density and bulk velocity of shy-B (Li et al. 2020). Matrosova et al. (2019) have shown that Ti-bearing shy-B with a *Pnnm* space group has a wider stability field than the Mg-Si end-member. Fe-bearing shy-B has much higher bulk modulus than the Fe-free end-member (Shieh et al. 2000; Li et al. 2016). The thermodynamic stability of Al-bearing shy-B expands to higher temperature with increasing Al content, such that shy-B with 31.9 wt% Al_2O_3 is stable up to 24 GPa and 2000 ˚C (Kakizawa et al. 2018). Based on the Al content and different substitution mechanisms in shy-B, two different types of shy-B were proposed by Kakizawa et al (2018). In a recent study, Kakizawa et al (2021) investigated an Al-bearing shy B sample with 1.9 Al pfu by FTIR spectroscopy, difference Fourier maps and bond-distance considerations. The sample contained more than the ideal hydrogen amount and they propose as additional H incorporation mechanism the complex substitution $2Mg^{2+} + Si^{4+} = 2\ Al^{3+} + 2\ H^{+} + \square^{Mg2+}$($\square^{Mg2+}$ means vacancy in Mg site). This needs further proof.

Knowing the elastic properties of shy-B is also very important for interpreting seismic anomalies in the deep transition zone and the shallow lower mantle, such as the low velocity zone, the 720 km discontinuity and high anisotropy zones (Rosa et al. 2015; Li et al. 2016; Yang et al. 2017). Thus, some experimental and *ab initio* computational studies investigated the elasticity on the Mg end-member of shy-B with different water contents at high P-T, and the results were used to explain seismic anomalies in the topmost lower mantle (Pacalo and Weidner 1996; Mookherjee and Tsuchiya 2015; Rosa et al. 2015; Li et al. 2016; Yang et al. 2017). However, Al-bearing shy-B showing a much wider P-T stability should be considered as an even more important candidate for H_2O transport and storage in the transition zone and topmost part of the lower mantle. One way to assess the hypothesis of Al-bearing shy-

125B phase being present in deeply subducted slabs is comparing mineralogical models
126with geophysical observations such as those from seismology. Here we provide
127information on the elasticity of shy-B as a function of the Al content.
128In this study we used rotating multi-anvil apparatus to synthesize a nominally Al-free
129shy-B with 0.04(1) Al atoms pfu (based on 18 O atoms pfu) and two Al-bearing shy-B
130with 0.47(2) and 1.35(4) Al atoms pfu. The syntheses products were analyzed with an
131electron microprobe (EMP), FTIR- and Raman-spectroscopy. The crystal structures of
132the three compositions were determined by single-crystal X-ray diffraction (XRD).
133Finally, the elasticity of Al-rich shy-B (with 1.35 Al atoms pfu) was investigated by
134Brillouin scattering (BS) at ambient conditions.
135

136**2. Experimental and analytical methods**

137Syntheses: Multi-Anvil apparatus

138Shy-B with three different compositions were synthesized in a rotating multi-anvil
139press at the GFZ, Potsdam (Deon et al. 2011). This device, featuring a Walker type
140high pressure module (Walker et al. 1990), is designed to allow continuous 360°
141rotation of the entire press at a maximum speed of 5° per second for the duration of
142the synthesis. Continuous rotation enhances a homogeneous distribution of fluid and
143solid within the sample capsule (see Schmidt and Ulmer 2004; Deon et al. 2011). We
144used a 10/5 assembly (lengths of the octahedral edges and the truncations of the WC
145cubes, respectively) with a MgO-based octahedron serving as pressure transmitting
146medium, a stepped graphite heater to avoid or minimize temperature gradients, and
147pyrophyllite gaskets. We monitored the sample temperature by using type C
148thermocouples (W5%Re-W26%Re). Details of the experimental setup are given in
149Mrosko et al. (2015). The starting materials were loaded in Pt capsules (length 2.25

150mm, outer diameter 1.4 mm), which were cold sealed. Experimental conditions, run
151number and the starting materials are summarized in Table 1.
152
153Electron microprobe analysis (EMPA)
154Several single crystals of the synthesized shy-B were handpicked, embedded in
155epoxy, polished and carbon coated in order to perform chemical EMPA. The analyses
156were performed using a JEOL Superprobe JXA-8230 (GFZ, Potsdam) in wavelength-
157dispersive X-ray diffraction (WDS) mode with the following crystal-analyzers: TAP
158for Si and Al, LIFL for Fe and TAPH for Mg. Although the samples do not contain
159Fe, Fe was required in the calibration and quality control process as the high-pressure
160ringwoodite standard contained iron. The microprobe was operated at 15 kV, 10 nA.
161The size of the measuring spot was 5 μm to prevent mineral destruction. Counting
162times on the peaks for Si, Al and Mg were 10 s and 20 s for Fe. Background counting
163times were 5 s for Si, Al and Mg and 10 s for Fe. The standard for Mg and Si was well
164characterized synthetic ringwoodite (MA-313, Mrosko et al. 2013), corundum for Al
165and hematite for Fe.
166
167FTIR spectroscopy
168FTIR spectra were measured using a VERTEX 80v FTIR spectrometer (Bruker
169optics) with an attached Hyperion II microscope at the IR-spectroscopy laboratory of
170the GFZ in Potsdam. A tungsten light source, InSb detector and a CaF_2 beamsplitter
171were used to measure the OH stretching range (2500 - 4000 cm^{-1}). Due to the high
172water content (the stoichiometric Mg-endmember has 5.8 wt.% H_2O) a very thin film
173of the sample material was required which was prepared from very small amounts of
174the fine grained portion of the syntheses product by pressing the material in a

diamond anvil cell (DAC) equipped with type II diamonds and without a gasket (thin-film preparation). The spectra were collected with the thin-films attached to one of the diamonds anvils and with 2 cm^{-1} resolution and averaged over 256 scans. The reference spectra were taken through the diamonds. The thickness of the first sample (MA-575) flattened in a DAC was determined by reflecting a broadband optical probe in the collected signal scaled by the refractive index of the sample. The refractive index was estimated to be n = 1.686 which is the average value along the three crystallographic axes (Pacalo and Weidner 1996). The thickness of MA-575 is 1.0 (1) μm. By optical observation we estimated the thickness of the thin films of MA-399 and MA-576 to be 1.0 and 1.3 μm, respectively. In the absence of a mineral-specific absorption coefficient, ε, we used the wavenumber-dependent method proposed by Libowitzky and Rossman (1997) to calculate molar absorption coefficients ε's for the bands of each sample. The spectra were deconvoluted with the software PeakFit by Jandel Scientific. We applied a Gaussian plus Lorenzian peak shape to all component bands. The integral absorbance (A_{int}) of each peak was summed up and multiplied by 3 to get the total integral absorbance $A_{int,tot}$. Density was derived from our single crystal X-ray refinements, which are 3.227(3), 3.179(3) and 3.166(4) g/cm^3 for MA-399, 575 and 576, respectively. We used the Beer-Lambert law to calculate the water content which can be found in Libowitzky and Rossman (1997).

Temperature-dependent thin-film spectra of samples MA-575 and MA-576 were also measured in the OH stretching region using the spectrometer and components described above in a Linkam FTIR600 cooling/heating stage adjusted to the microscope. Spectra were collected with a spectral resolution of 2 cm^{-1} and averaged over 1024 scans. Shy- B thin-films were investigated from ambient conditions down to -180 °C. To reach the low temperature the stage was cooled with liquid nitrogen.

The low-T spectra of sample MA-575 exhibited an additional peak at around 3180 cm^{-1} at -180 °C, which we attributed to the formation of ice crystals, visible to our eyes. In experiment MA-576 we avoided the formation of ice by annealing the sample in the Linkam sample chamber for 1 h at 100 °C while purging with nitrogen gas before the cooling experiment started.

To verify the presence of weak bands hidden in the background of the thin-film spectra also spectra of large crystals (about 100 x 50 μm^2) with thicknesses varying between 30 to 60 μm were measured in the OH stretching region using the spectrometer and components described above. The crystals were placed on KBr sample holders and spectra were averaged over 128 scans.

FTIR spectra in the lattice vibrational range (400-2000 cm^{-1}) were measured using the spectrometer described above but with a Globar light source, MCT detector and a KBr beamsplitter. The samples were prepared as thin films (see above) and placed on KBr sample holders. The spectra were collected with 4 cm^{-1} resolution and averaged over 1024 scans.

Raman-spectroscopy

Raman spectra were measured with a HORIBA Jobin Yvon LabRAM HR800 VIS spectrometer (GFZ, Potsdam) equipped with a green 514 nm wavelength diode-pumped solid-state laser. The spectra were collected from the microprobe mounts after removal of the carbon coating.

Spectra in the range of 100–1000 cm^{-1} were measured for phase identification and the data acquisition time was 30 s for MA-399 and MA-575 and 50 s for MA-576. We measured several different single crystals for MA-576 - all gave much weaker signals than the crystals of the other two samples. The spectra were deconvoluted with the

225software PeakFit by Jandel Scientific.

226Spectra in the OH stretching region were measured in the range of 3200 – 3600 cm^{-1} 227with the same configuration. These measurements were extremely difficult as the 228samples were very sensitive to the laser light decomposition and showed high 229fluorescence. The Al-bearing samples were measured with 50 % laser power. Spectra 230of MA-575 were collected for 10 sec with 3 accumulations while spectra of MA-576 231were measured for 20 sec with 3 accumulations.

232

233Single-crystal X-ray diffraction

234Single-crystal XRD data were collected at the Extreme Conditions Beamline (ECB, 235P02.2) at PETRA III, DESY (Hamburg, Germany). In order to improve data quality 236and for the purpose of absorption correction, we pre-selected samples of a shape 237similar to a sphere. Samples were glued on the top of a tungsten needle (2 μm tip), 238which acted as holder for the XRD measurements, enabling data collection during 239rotation (-90° - +45°) around the vertical axis. The energy of the beam was tuned to 24042.7 keV, while the beam was focused to 3 x 8 μm^2 (H x V, full width at half 241maximum). Sample to detector distance (SDD) and tilt of the detector were initially 242determined from the powder X-ray diffraction image of a CeO_2 standard from NIST 243(674b) using DIOPTAS (Prescher and Prakapenka 2015) and further refined in 244CrysAlisPro© from Rigaku. The single-crystal data are calibrated by an enstatite 245crystal. Crystal structures were solved using OLEX2 with a SHELX backend 246(Dolomanov et al. 2009; Sheldrick 2015). Additional data analysis was conducted 247with JANA2006 (Petříček et al. 2014). The structural models displayed below were 248drawn using the software VESTA (Momma and Izumi 2011).

249

Brillouin Scattering

We collected single-crystal Brillouin scattering spectra from grains selected from sample MA-576, our Al-bearing shy-B sample that contained the highest amount of Al (1.35 atoms pfu). Four crystal platelets were double side polished to 20-25 μm final thickness, and their orientations were determined by single crystal XRD at the ECB (see above). The orientations of the four platelets were (-0.5 , 0 , 0.87), (0.83 , -0.56 , -0.08), (-0.78 , -0.38 , 0.49) and (-0.81 , -0.59 , 0.09). Brillouin scattering experiments at ambient condition were performed at the laser lab of the Extreme Conditions Science Infrastructure (ECSI) of PETRA III, DESY, Hamburg, established by the GFZ Potsdam. The signal was excited by a Nd:YVO_4 laser with a wavelength of 532 nm and an output power of 100 mW. The Brillouin spectra were recorded by a six-pass Sandercock tandem Fabry-Perot interferometer. The external scattering angle (i.e. the angle between incident and scattered light directions outside the sample) was 50°, and the mirror spacing was 6 mm during the measurements. The acoustic velocities of the sample were calculated from the measured Brillouin frequency shift, $\Delta\upsilon_B$, as:

$$v=\frac{\Delta\upsilon_B\lambda_0}{2\sin(¿\theta/2)¿} \quad (1)$$

where v is the acoustic velocity, λ_0 is the laser wavelength of 532 nm, and θ is the external scattering angle. The dataset of individual acoustic velocities, and corresponding wave vector orientations and polarizations was used to determine the 9 independent nonzero elastic coefficients of the Al-shy-B sample. The elastic constants were determined by inverting a set of Christoffel's equations:

$$\left|C_{ijkl}n_jn_l-\rho v^2\delta_{ik}\right|=0 \quad (2)$$

where C_{ijkl} is the elastic constants, n_j and n_l are the direction cosines of the acoustic

wave propagation direction. ρ is the density, v is the acoustic velocity derived from Brillouin frequency shift, and δ_{ik} is the Kronecker delta. Based on the volume derived from X-ray diffraction data and the composition $Mg_{8.04}Si_{2.17}Al_{1.35}O_{18}H_{7.18}$, the density of MA-576 was calculated to be 3.166 (4) g/cm^3, the number in parentheses is 1σ standard deviation on the last digit.

3. Results

Table 2 shows the EMPA results and chemical composition for the three different syntheses. Sample MA-399 contains 0.35(3) wt.% Al_2O_3, which corresponds to 0.04 atoms per formula unit in shy-B. This amount has negligible structural effects, and we will consider this sample as Al-free in our discussion. The two samples MA-575, MA-576 contain 4.05(18) and 11.72(21) wt.% Al_2O_3, respectively (Table 2). The H_2O contents of the 3 samples were determined by FTIR absorption, and are equal to 8(1), 10(1), 11(1) wt.%, respectively. To calculate the chemical formula, we preferred to take the IR-determined water contents, rather than the difference of the EMP sum to 100 %. The complete chemical formulas for each synthesis are reported in Table 2. Our results show that water content increases with Al content, which is consistent with the result for shy-B with moderate Al content (0.85<Al<1.91 pfu) in Kakizawa et al (2018). There are minor amounts of ringwoodite coexisting with shy-B in MA-399 and pyrope coexisting with Al-bearing shy-B in MA-575 and MA-576, which means that pyrope will coexist with Al-bearing shy-B in Al-rich systems.

In our syntheses, we did not observe complex phase assemblages as reported by Kakizawa et al (2018). In particular, we did not observe a coexistence of two shy-B phases with different Al content in the same synthesis run. We contribute this due to the rotating multi-anvil press setup, used in our experiments. The rotation of our

multi-anvil press during the duration of the experiments ensures homogeneous distribution of the liquid phase in the capsule. Consequently, all the chemical compositions of different grains of our sample are nearly identical within the uncertainty of our measurements. We did not observe other coexisting high temperature phases such as δ-AlOOH (Kakizawa et al. 2018) as the temperature of our syntheses is 200-400 ˚C lower than those of Kakizawa et al. (2018).

Figure 1 shows the FTIR spectra in the OH stretching region of the samples prepared as thin-films. In the spectrum of the nominally Al-free shy-B sample (MA-399) we can distinguish two main OH bands (ν_1 and ν_2), which we assign to the vibrations of two different hydroxyl groups according to space group symmetry *Pnn*2 that are consistent with Koch-Müller et al. (2005, 2014) and Trots et al. (2015). With increasing Al-content the frequencies of the OH bands do not change. However, the bands visibly broaden, while at least a third OH band ν_3 is visible at lower wavenumbers. The intensity of ν_3 increases with increasing Al-content (Figs.1, S1). In the FTIR spectrum of the sample with the highest Al-content (MA-576) only two very broad OH bands can be distinguished (merging $\nu_1 + \nu_2$ and additional ν_3). In all spectra (Fig. S1) a fourth very weak OH band can be identified at higher wavenumbers, e.g. at 3554 cm^{-1} (MA-576). The peak positions, the integral absorbances normalized to 1 cm and the band widths (FWHM) determined for the thin-film spectra are listed in Table S1. The calculated water contents are listed in Table 2. All samples contain more OH than the ideal value pointing to two hydrogen positions. As also indicated by the FTIR spectra there must be additional hydrogens incorporated in the crystal structure. Low temperature spectra of sample MA-575 and MA-576 were collected in the hope to reduce the FWHM of the OH bands for a better band assignment. But as shown in Fig. S2 there is nearly no change in the spectra with decreasing temperature.

To finally verify the presence of ν_3 and ν_4 in all three samples we collected spectra of individual crystals in the OH stretching region. The spectra are shown in Fig. S1. The spectra are oversaturated in the range of 3450 to 3300 cm-1 but the presence of the ν_3 band at lower wavenumbers can clearly be seen also in MA-399. In all three spectra the very weak band ν_4 at higher wavenumber is clearly present.

FTIR spectra in the range of the lattice vibrations are shown in Fig. 2 and their frequencies are reported in Table S2. Based on the results of Hofmeister (1999), we assign specific vibrations to the peaks; some peaks remain unassigned (Table S3). The spectra of the three samples are very similar - they are dominated by vibrations of tetrahedrally and octahedrally coordinated Si. The main difference between the Al-free sample MA-399 and the Al-bearing sample MA-576 is in the region around 600 – 700 cm^{-1} and we tentatively assigned the two most prominent bands in the spectrum of the sample MA-576 to vibrations of octahedrally coordinated Al.

Raman spectra in the range of the lattice vibrations are shown in Fig. 3. Nine Raman peaks for each Shy-B were fitted, and their frequencies are reported in Table S2. Based on the results of Hofmeister et al. (1999), we assign specific vibrations to the peaks; some peaks remain unassigned (Table S2). The Raman spectra show that the intensity of the peaks becomes weaker with increasing Al-content (Fig. 3). Similar behavior has been reported that Al-rich phase D shows broader and weaker Raman peaks when compared with Al-free one (Pamato et al. 2015; Xu and Inoue 2019). The values of FWHM show that most peaks broaden with increasing Al content (Table 2), potentially indicating presence of cationic disorder. Most of the peak positions move to larger wavenumbers as a function of increasing Al content. This observation indicates that the incorporation of the Al increases the vibrational frequencies, and is compatible with the observed decrease of the unit cell volume (see below). Raman

spectra in the OH stretching region also show increasing peak broadening with increasing Al incorporation (Fig. S3).

In order to examine the structure of shy-B crystals with different Al content, we performed single-crystal X-ray diffraction refinements at ambient condition. The details can be found in the cif files in supplementary materials. The unit cell parameters are also listed in Table S3. Although our FTIR results clearly indicate two different hydrogen sites suggesting the space group *Pnn*2 for nominally Al-free shy-B (MA-399), we use both *Pnnm* and *Pnn*2 during the refinement to examine all the possibilities for the other two compositions (MA-575 and MA-576). The main difference between *Pnnm* and *Pnn*2 structural models for shy-B is the description of H atom positions. *Pnnm* has just one general H position (multiplicity 8) while *Pnn*2 has two general H positions (multiplicity 4). Considering similarity of both space groups and lattice parameters, we can compare them for shy-B with different Al contents (Fig.4). Our single-crystal refinement shows that the *Pnnm* and *Pnn*2 offer the same cell parameters. Figure 4 shows that the incorporation of 1 Al pfu will reduce *a*, *b*, *c* unit cell parameters and the volume (*V*) of Al-bearing shy-B by 0.69%, 0.24%, 0.16% and 0.29%, respectively.

Apart from the lattice parameters, the incorporation of Al affects various structural features, including the volumes of MgO_6, SiO_4 and SiO_6 polyhedra. Based on the results of our analysis, we can also compare the volume of the different MgO_6 octahedra as a function of Al content (Fig. 5; Fig. S4). In order to simplify the discussion, it seems justified to continue using *Pnnm* as a reference. Indeed, the octahedral sites of Mg5, Mg6, Mg7 in *Pnn*2 are structurally related to those of Mg3, Mg4, Mg1 in *Pnnm*. In fact, the solutions for the both space groups are very close in terms of atomic positions and interatomic distances. This is not surprising, given that

374the difference between the space groups arises from the symmetry break due to
375position of H, the atom with the lowest Z, which weakly contribute to X-ray
376diffraction.

377As the Al content in shy-B increases, the volume of the Mg1 and Mg2 octahedra are
378reduced, in contrast to a slightly increased volume of Mg4 octahedron. The changes of
379Mg3 octahedron volume are insignificant (Fig.5). The sites containing Si behave
380differently as Al content changes, i.e. the volume is increased with increasing Al
381content for SiO_6 octahedra but remains unchanged for SiO_4 tetrahedra (Fig.5)
382indicating low probability of Al substitution at the latter site.

383In order to put strong constraints on elasticity, we collected Brillouin scattering
384spectra and determined acoustic velocities from four crystal platelets, even though
385two platelets have similar orientations (Fig.6). Due to the extensive dataset, the
386inversion to determine the full elastic tensor yielded unambiguous results. The fitting
387results are shown in Figure 6 and the elastic stiffness coefficients are listed in Table 3.
388Al-bearing shy-B shows lower C_{11}, C_{44}, C_{55}, C_{13}, C_{66} but higher C_{22}, C_{12} than Al-free
389end member and similar C_{33}, C_{23}. The adiabatic bulk modulus (K_S) and the shear
390modulus (G) were estimated to be 140.7(7) GPa and 88.2(4) GPa, respectively, which
391are also lower than the Al-free composition. The azimuthal V_P anisotropy and
392maximum V_S^{PO} splitting are calculated using the following equations

393

394$$A_P = \frac{V_{Pmax} - V_{Pmin}}{V_{Pavg}} \times 100$$

395$$A_S^{PO} = \frac{Max(V_{S2} - V_{S1})}{V_{Savg}} \times 100$$

396where V_{Pmax}, V_{Pmin}, V_{Pavg}, are maximum, minimum and average V_P in the crystal
397structure. Max (V_{S2} - V_{S1}), describes the maximum difference between V_{S2} and V_{S1} in

the same propagation direction; V_{Savg} computes the average of all the V_S. The values of A_P and A_S^{PO} are 20.1% and 15.9%, respectively.

4. Discussion

4.1 Incorporation of H and Al in the structure of shy-B

Based on our single-crystal X-ray refinements of shy-B with different Al contents, we can infer the mechanisms of Al incorporation in the shy-B structure based on the volume of MgO_6 and SiO_6 octahedra. In the structure of Al_2O_3 (corundum structure), MgO (periclase structure) and SiO_2 (rutile structure), used as simple reference, the volume of AlO_6 octahedra is larger than that of the SiO_6 octahedra, but smaller than the MgO_6 octahedra (Fig. S5). Based on this simple comparison we expect as a first order approximation that the incorporation of Al by substitution of Si in octahedral sites will increase the volume of the octahedra, while it will lower the volume of the octahedra when it substitutes Mg. We observe that with increasing Al content in shy-B, the volume of the SiO_4 tetrahedral site remains unchanged and that of the SiO_6 octahedron becomes larger (Fig. 5) indicating that Al^{3+} substitutes for the octahedral Si^{4+} which is consistent with the Al substitution in stishovite (Bromiley et al. 2006; Litasov et al. 2007a). The volumes of Mg1 and Mg2 octahedral sites are decreasing with increasing Al content, indicating a preferred Al substitution for Mg in Mg1 and Mg2 (if we refer to the *Pnnm* space group). The volume of the Mg3 octahedron remains unaffected by the change in Al content suggesting that the substitution of Al in Mg3 octahedra site is negligible. In contrast, the volume of the Mg4 octahedral site slightly increases with increasing Al content, indicating a presence of Al in Mg4 octahedral site to should be limited. The increase of Mg4 octahedron might be caused by vacancies in this site.

The hydrogen positions and OH dipole directions of the stoichiometrically incorporated hydrogens are well known from neutron diffractions data and polarized FTIR spectra on oriented crystals (Trots et al. 2013, Koch-Müller et al. 2014); the two main OH bands in MA-399 are assigned to these two OH dipoles, and their vibrations are labelled as ν_1 and ν_2. In the thin-film spectra there may be hints for two additional OH bands ν_3 and ν_4. To have a better insight into the defect H we collected low temperature spectra and spectra on relatively thick crystals (Figs. S1, S2). There is nearly no change in the OH bands with decreasing temperatures. The bands remain broad, indicating that the FWHM is caused by cationic disorder rather than disorder of by thermal broadening. According to Trots et al. (2013) both non-defect hydrogens are coordinated by Mg1, Mg2 and Mg3 sites. We suggest that Al-substituting Mg in Mg1 and Mg2, as shown in this study, leads to variable environments of the hydrogen and thus to the observed higher band widths. Fig. 1 and S1 clearly shows that defect hydrogen is present in all three samples. The intensity of the band ν_3 at 3261 cm^{-1} (MA-399) increases with increasing vacancies of the Mg sites for the samples MA-575 and MA-576 and thus we assigned the band to this defect. The band ν_4 is also present in all three sample, however, only with a constant and minor intensity. We will discuss its possible origin later on.

Due to the very limited Al content in MA-399 whose composition is $(Mg_{9.39}\square_{0.61})(Si_{2.93}Al_{0.04}\square_{0.03})H_{5.37}O_{18}$, we considered it as Al-free shy-B in this discussion. The extra 1.37 H pfu goes to the tetrahedral silicon site as hydrogarnet substitution (0.12 H pfu), vacancies in the Mg site (1.22 H pfu) and neglectable 0.04 H pfu goes to Si octahedral site by the mechanism of $Si^{4+}=Al^{3+} + H^{+}$.

Based on our single-crystal structural models and electron microprobe analyses of the Al-bearing samples MA-575 and MA-576, we propose the following coupled

448substitutions for Al-bearing shy-B: (1) $Si^{4+}=Al^{3+} + H^{+}$; (2) $2\ Al^{3+} = Mg^{2+} + Si^{4+}$; (3)
449$\square^{Mg2+} = 2H^{+}$. As the ionic size of $^{VI}Al^{3+}$ and $^{VI}Si^{4+}$ are quite similar (Shannon, 1976),
450we propose that Al^{3+} first occupies all the Si octahedral sites and then the other sites.
451We rule out the possibility of large contributions from the hydrogarnet substitution as
452the X-ray refinement indicate fully occupied (Si + Al) tetrahedral and octahedral Si
453sites. However, a limited amount as observed in MA-399 may be possible.
454Sample MA-575 contains 0.47 Al pfu in total ($Mg_{8.75}Al_{0.47}Si_{2.62}H_{6.60}O_{18}$). 0.38 Al pfu is
455needed to occupy all the Si octahedral sites, and the remaining 0.09 Al pfu goes to the
456Mg site. We thus assign 0.29 Al to the octahedral Si-site via mechanism (1); 0.09 Al
457pfu (in total) goes to Mg1 and Mg2 sites and 0.09 pfu Al^{3+} to the octahedral Si site by
458mechanism (2). Therefore, the residual charge in the Si site should be balanced via
459mechanism (1) by 0.29 H pfu. Vacancies in the Mg sites are charge balanced by
460mechanism (3) resulting in 2.32 H pfu. The crystal-chemical formula for MA-575 can
461thus be stated as $(Mg_{8.75}Al_{0.09}\square_{1.16})\ (Si_{2.62}Al_{0.38})\ H_{6.60}O_{18}$.
462Following the substitution mechanisms as described for MA-575, the crystal-chemical
463formula of MA-576 will be $(Mg_{8.04}Al_{0.52}\square_{1.44})(Si_{2.17}Al_{0.83})H_{7.18}\ O_{18}$: 0.31 Al pfu are
464assigned to the octahedral Si site via mechanism (1) results in 0.31 H^{+} pfu; 0.52 Al
465pfu (in total) is in the Mg1 and Mg2 sites and 0.52 Al pfu in the octahedral Si site via
466mechanism (2); vacancies in the Mg sites are charge balanced by mechanism (3)
467resulting in 2.88 additional H pfu.
468In summary, with increasing Al-incorporation in shy B the Mg vacancy content and
469Al incorporation into the octahedral Si site increase, and both lead to an increasing H
470content.
471The assignment is consistent with the FTIR spectra: the increasing absorbance of the
472band ν_3 with increasing Al- content is due to increasing Mg vacancies content (0.6,

473 1.16 and 1.44 pfu, in MA-399, MA-575 and MA-576 respectively), and charge
474 balance via mechanism (2) resulting in 1.4, 2.32 and 2.88 H^+ pfu. In the nominally Al-
475 free sample MA-399 the absorption band ν_3 is weak but still clearly visible (Fig. S1;
476 see also Table S1). There should be additional weak OH bands due to mechanism (1)
477 in the FTIR spectra of MA-575 (0.29 H) and MA-576 (0.31 H) but nearly none (0.04
478 H pfu) in the spectra of MA-399. These additional OH bands should have frequencies
479 close to 3100 cm^{-1} based on available data for H-bearing Al-stishovite (Bromiley et al.
480 2006; Litasov et al. 2007a) and they may be hidden in the background. However, we
481 do observe a weak OH band, ν_4, in all three samples and we tentatively assign it to the
482 hydrogarnet substitution. At least for sample MA-399 the chemical composition
483 analyses indicate 0.03 pfu vacancies in the Si site which would result in 0.12 H pfu.
484 Due to the complex Al incorporation into the structure and variable vacancies we
485 cannot rule out that the same holds for MA-575 and MA-576. The hydrogarnet
486 substitution would result in four OH^- bands with a strong band in the frequency range
487 of ν_4 – the others may be hidden by the high intrinsic OH described above.
488 Our interpretation of the data is partly consistent with the results of Kakizawa et al.
489 (2021). They investigated one Al-shy B sample with nearly the same composition as
490 MA-576. The FTIR spectra are quite similar to the spectra of MA-576 shown in Fig. 1
491 and S1. In agreement with our conclusions the broad band at 3400 cm^{-1} (our merged
492 v_1, v_2) is assigned to the stoichiometrically incorporated hydrogen for *Pnn*2. The band
493 at 3260 cm^{-1} (our band v_3) and at 3540 cm^{-1} (our band v_4) are according to Kakizawa et
494 al. (2021) assigned to the substitution mechanism (1) ($Si^{4+} = Al^{3+} + H^+$) and to a new
495 and complex substitution mechanism (4): 2 $Mg^{2+} = Al^{3+} + H^+ + \square^{Mg2+}$, respectively.
496 This assignment is not supported by our data. We studied three samples with different
497 Al-contents, the band in question ($v3$) is present in all three spectra (Fig. S1) but only

498two of the samples show the substitution of Si^{4+} by Al^{3+} (Table 2). In addition, in our
499study the band at 3450 cm^{-1} (ν_4) shows a constant and weak absorption in the spectra
500of all three samples (Fig. S1), which is not consistent with the substitution mechanism
501in which Mg vacancies are involved as the amount of vacancies increases with Al
502content.

503Al incorporation into the Mg-site (mechanism 2) causes a broadening of the OH
504absorption bands, which leads to the merging of OH bands ν_1 and ν_2. As observed in
505other solid solutions systems such as Al-bearing phase D, the partial substitution of
506Mg by Al at the Mg site results in variable Mg-Al environments for the H sites and
507causes band broadening (Xu et al.,2019). We can rule out that the band broadening is
508due to dynamic disordering of H atom around the normal H position as we see no
509sharpening of the bands with decreasing temperature (fig s2). In conclusion, Al-free
510shy-B with 2 distinct H atomic position crystallizes in *Pnn*2 space group and the
511incorporation of Al into the hydrogen coordinating Mg sites M1 and M2 leads to
512cationic disorder and seems to result in overlapping unresolved OH bands leaving the
513question of the ultimate structure type (*Pnn*2 or *Pnnm*) for Al-rich shy-B unsolved. To
514gain further structural information, Al-rich shy-B needs to be examined by NMR or
515neutron diffraction in future.

516

5174.2 Elasticity of Al-bearing shy-B

518The elasticity of Al-free shy-B with different hydrogen contents has been investigated
519at ambient or high P-T conditions using Brillouin scattering in 3 studies (Pacalo and
520Weidner 1996; Rosa et al. 2015; Li et al. 2016). We list the sets of 9 independent
521elastic constants (at ambient conditions) for the different chemical compositions in
522Table 3 and we plot them in Fig.7 where the valency of the cations is considered (i.e.

for MA-576 (Al+H)/(H+Al+Mg+Si) and equates to (3*1.35 + 7.18)/36). We also recalculated the results of Rosa et al (2015) using the correct composition (and density) of that sample material following the suggestion by Li et al (2016). The elastic stiffness coefficients of Al-free shy-B show a general decrease with increasing H content, and only C_{44} and C_{12} are insensitive to the H content within uncertainty; the effect of Al incorporation modifies these trends, especially in the case of C_{11}, C_{22} and C_{23} (Fig.7). The incorporation of H in shy-B will significantly reduce the mass per formula unit leading to a sensible decrease of density of shy-B. To better identify and discuss the effects of H and Al incorporation on the elasticity and acoustic velocity of shy-B, we try to exclude the effect of density (mostly due to H incorporation) by normalizing the individual coefficients and expressing them as the ratio C_{ij}/density. Using the value of C_{ij}/density from Pacalo and Weidner (1996) as a reference ideal Mg- end member of shy-B, we calculated the differences of the individual C_{ij} for each composition of shy-B (Fig.8). When the effect of density (i.e. of hydrogen incorporation) is removed, the value of most C_{ij}/density for Al-free shy-B with different water contents remains equal within uncertainties (Fig. 8). The differences of C_{12}, C_{13} and C_{23} are larger, as the off-diagonal elastic coefficients are more difficult to constrain (they cannot be determined independently by individual Brillouin scattering or sound velocity measurements). When we consider the Al-free system in previous studies, the values of C_{ij}/density for all the other coefficients are effectively insensitive to the H content indicating that the incorporation of H in excess compared to the ideal Mg- end member composition will not affect the sound velocity of shy-B, because it equally reduces both the elastic coefficients and the density. When we consider Al-bearing shy-B (MA-576), the elastic coefficients C_{11}, C_{44}, C_{55}, C_{66} and C_{13} are lower, C_{22}, C_{12}, C_{23} are higher and C_{33} and C_{66} are equal to Al-free shy-B. The

large difference observed for several density-normalized C_{ij} of Al-bearing shy-B indicates that the substitution of Al in the structure has a much larger impact on sound velocity of this phase.

In order to understand the relationship between Al incorporation and elasticity, we need to consider the crystal structure of Al-free shy-B. The structure of shy-B is characterized by an alternation of layers containing octahedra (Mg) and tetrahedra (Si) and layers containing octahedra (Mg) and octahedra (Si) stacked along the *b*-crystallographic axis (Fig.9). All the available experimental studies of the elastic properties of Al-free shy-B (with variable H content) and also density functional theory (DFT) studies consistently show that *b* is the stiffest axial direction with C_{22} ~10% larger than C_{11} and ~7% larger than C_{33} (Pacalo and Weidner 1996; Mookherjee and Tsuchiya 2015; Rosa et al. 2015; Li et al. 2016; Yang et al. 2017). The stiffness along the *b*- direction is enhanced by the presence of two different alignments or columns of polyhedral units running along this direction, and it is controlled by the stiffest unit in these alignments. Based on the available data the stiffness of Si-O tetrahedra is much higher than that of Mg-O and Si-O octahedra (Hazen and Finger 1982). In both columns aligned along *b,* one of the stiff Si-O tetrahedral bonds is systematically aligned parallel to the *b*-axis (Fig. 9). In each column, both Si-tetrahedra and softer Si-octahedra are present with a ratio of 1:1. However, due to a shift along *b*-axis between the two columns, the resulting sequence of layers stacked along *b* have a ratio of tetrahedra-bearing vs. fully octahedral layers of 2:1. The distance between Si atoms aligned along the *b*-columns is a sequence of ~3.5 Å and ~5.2 Å.

Looking for structural features parallel to the *a*- and *c*-axes, we find that 2 types of column-like units have either only Si-octahedra or Si-tetrahedra alternating with Mg-

573octahedra. The distance between Si atoms along the columns running parallel to *a* is
574equal to the unit-cell parameter *a* (~5 Å). No stiff bond is aligned along the *a*-
575direction, and the arrangement of the stiff tetrahedra gives the option for rigid
576rotations to accommodate compression. Column-like structures along the *c*-axis offer
577a more direct alignment of cations, involving either only Si-octahedra or Si-tetrahedra,
578at distances equal to the *c*- unit-cell parameter (~8.5 Å). However, the packing of
579polyhedra is denser than along the *a*-axis, with average distances between Si and Mg
580or Mg and Mg averaging at ~2.7 Å.
581The mechanisms through which Al incorporation affects the elastic tensor of shy-B is
582complicated by the coexisting effects of vacancies and hydrogen incorporation.
583Below, we interpret the effects on the compressional coefficients C_{11}, C_{22}, and C_{33}
584based on our complete set of structural and spectroscopic observations.
585We identify major mechanisms: (1) Al incorporation causes a decrease of the unit cell
586parameters *a*, *b*, *c* and potentially stiffening all the C_{ii} (i=1,2,3); (2) Al substitution at
587the octahedral Si site softens Si-O octahedra; (3) Al substitution at Mg sites stiffens
588Mg-O octahedra; (4) vacancies on Mg octahedra make them softer. In the case of C_{11}
589(the uniaxial stiffness along the *a*- direction) the mechanisms (1), (2) and (3) are all
590active. However, as the distance between Si polyhedra is very large (~5 Å), the
591shrinking of *a*- axis has a limited effect on C_{11}. In addition, in MA-576 a large amount
592of Si is substituted by Al and mechanism (2) is dominant over (1) and (3).
593Consequently, C_{11} softens with respect to Al-free shy-B. In the case of C_{22}, as the
594stiffest Si-O tetrahedral aligned with *b*-axis, mechanisms (1) will dominate and make
595C_{22} stiffer. In the case of C_{33} the four different mechanisms balance each other, and the
596value of the constant is only marginally affected by Al incorporation.
597

5. Implications

Shy-B is potentially a major hydrous phase in cold slabs subducted to the depth of the transition zone and the topmost part of the lower mantle (Schmidt and Ulmer 2004; Ohtani 2005). The presence and decomposition of Al-free shy-B might be related to the low velocity zone or 720-800 km discontinuity in the shallow lower mantle (Li et al. 2016; Yang et al. 2017). The study by Kakizawa et al. (2018) showed that Al-bearing shy-B has a much wider stability field compared to the Al-free counterpart. Pyrope-rich garnet can be the main Al carrier in the subducted slabs in the deep transition zone and the top of the lower mantle. Our syntheses results show that Al-free shy-B coexists with ringwoodite, while Al-bearing shy-B coexists with pyrope at the conditions of our synthesis runs (Table 2). In the peridotite layer of subducted slabs, Al-free shy-B could be favored compared to Al-bearing shy-B because the amount of available Al is low. On the other hand, in the basaltic layer, where pyrope is the most abundant phase, Al-bearing shy-B would be present coexisting with pyrope-rich garnet if the subducted slab is H_2O rich. The Al-free shy-B decomposes around 30 GPa and 900 K indicating that it cannot represent a water carrier even at the conditions of a very cold subduction slab (Ohtani et al. 2003). Having a wider stability, Al-bearing shy-B with 31.9 wt.% Al_2O_3 could be stable at 25 GPa and 2300 K and could be a potential water carrier in the topmost lower mantle (Kakizawa et al.2018). The elastic properties of Al-bearing shy-B are important for our understanding of the low seismic velocity anomalies and anisotropy anomalies observed in the transition zone and topmost part of the lower mantle. Our study demonstrates that while the incorporation of additional H_2O in shy-B mainly affects the density and has only a slight effect on sound velocity (Fig. 8 and Table 3), on the opposite the incorporation of Al has a more complex effect on density and elasticity

623resulting in a relevant effect on sound velocity and anisotropy. Comparing V_P and V_S
624with that of standard Al-free shy-B, Al-bearing shy-B has 1.8% and 2.6 % lower V_P
625and V_S, respectively (Fig. 10). A previous study of sound velocity of H_2O-rich Al-free
626shy-B at high P-T has shown that the presence of shy-B could help to explain the low
627velocity zone at the topmost lower mantle but cannot help to understanding high
628anisotropy regions (Li et al., 2016). Our Al-bearing shy-B displays densities that are
629comparable to those of the H_2O-rich Al-free shy-B investigated by Li et al (2016),
630which has similar H_2O content (Table 3.). At ambient conditions, Al-bearing shy-B
631has the lowest acoustic velocities among all the previously investigated shy-B
632compositions (Fig.10). More importantly, the V_P anisotropy and V_S^{PO} splitting of Al-
633bearing shy-B are 13.2% and 6.1% higher than those of Al-free one (Fig.10). The
634volume fraction of shy-B in the hydrated pyrolite system can reach 25% at pressures
635of the topmost lower mantle (Ohtani et al.2005). Assuming a similar pressure and
636temperature dependence of the elastic tensor as Al-free shy-B (Li et al., 2016), the
637presence of 25 vol.% of Al-bearing shy-B in deep subducted slabs may produce
638observable lower seismic velocity at topmost lower mantle. More interestingly, Al-
639bearing has higher elastic anisotropy than Al-free shy-B, we thus conclude that if 25
640vol.% of Al-bearing shy-B with strong lattice preferred orientation (LPO) might cause
641~2% anisotropy for V_S and help to understand the anomalous seismic features
642observed inside and in proximity of Tonga deep subducted slab (Chen and Brudzinski
6432003). However, our studies only focus on the elasticity at ambient condition, and a
644high P-T study investigating the elasticity of Al-bearing shy-B is necessary in the
645future.

646

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806

807Figure Captions

808**Figure 1.** Infrared absorption spectra of shy-B with different Al content in the
809frequency range of the OH stretching vibration.

Figure 2. Infrared absorption spectra of shy-B with different Al content in the frequency range of the lattice vibrations.

Figure 3. Raman spectra of shy-B with different Al content at the frequency range of 100-1000 cm^{-1}.

Figure 4. Unit cell parameters and unit cell volume of shy-B with different Al contents. Error bars are smaller than the symbols.

Figure 5. Polyhedral volumes of various Mg and Si sites for shy-B with different Al contents in *Pnnm* space group. (a) Dependence of the volumes of MgO_6 octahedra on the Al content. Squares, circles, triangles and diamonds represent the octahedral volume for the site of Mg1, Mg2, Mg3 and Mg4, respectively. (b) Dependence of the volumes of SiO_6 octahedron and SiO_4 tetrahedron on the Al content; triangles and circles represent the octahedron and tetrahedron of SiO_6 and SiO_4, respectively. Error bars are smaller than the symbols in both the panels.

Figure 6. Acoustic velocities of Shy-B with 1.35 Al pfu (sample MA-576) measured on four platelets with different orientations. Circles: quasi-longitudinal acoustic velocity; squares: fast quasi-shear acoustic velocity; triangles: slow quasi-shear acoustic velocity; curves: fitting results; black, red, green and blue colors represent the orientations of (-0.5 , 0 , 0.87), (0.83 , -0.56 , -0.08), (-0.78 , -0.38 , 0.49) and (-0.81 , -0.59 , 0.09), respectively.

835**Figure 7.** Relationship between elastic stiffness coefficients and the amount of H, Al
836and Fe incorporated in shy-B. The valency of the cations is considered (i.e. for MA-
837576 the (Al+H)/(H+Al+Mg+Si) is equal to (3*1.35 + 7.18)/36). The solid circles
838represent this study for Al-bearing shy-B. The open circles represent the Al-free shy-
839B with different water contents from previous studies (Pacalo and Weidner. 1996;
840Rosa et al. 2015; Li et al.2016); the open red diamonds in panel (d) represent the
841isothermal bulk modulus (K_{T0}) fitted to X-ray diffraction for Al-free shy-B with
842different water contents (Litasov et al. 2007b; Inoue et al. 2013); the open red star
843represents the K_{T0} of Fe-bearing shy-B (Crichton et al. 1999). Error bars smaller than
844the symbols are not shown in the figures.

845

846**Figure 8.** Effect of H and Al incorporation on the individual elastic stiffness
847coefficients of shy-B. The change is expressed as $100*[(C_{ij}/\text{density})_a - (C_{ij}/\text{density})_{ref}]$,
848where the subscript a indicates Al-bearing or H-enriched shy-B and the subscript ref
849indicates ideal Mg- end member shy-B (Pacalo and Weidner, 1996). Solid circles
850MA-576; open blue circles from Li et al.(2016); open green circles from Rosa et al.
851(2016)

852

853**Figure 9.** Crystal structure of Al-bearing shy-B(MA-576): *b*-*c* plane (panel (a)) and *a*-
854*c* plane (panel (b)).

855

856**Figure 10.** Average aggregate sound velocity and acoustic anisotropy of Al-bearing
857and Al-free shy-B at ambient condition. Solid circles: this study; open circles: Pacalo
858and Weidner (1996); Rosa et al. (2015); Li et al. (2016). (a) sound velocity of Al-
859bearing and Al-free shy-B. (b) V_P anisotropy and V_S^{PO} splitting of Al-bearing and Al-

860free shy-B.