New findings from old data: A semi-experimental value for the *eQq* of the nitrogen atom

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ABSTRACT

Nuclear quadrupole interaction arises from the coupling of the nuclear quadrupole moment with the electric field gradient. Thus, it is associated with electron filling and the electronic structure of molecules. We demonstrate a simple method for planar molecules based on a direct correlation between the out-of-plane quadrupole coupling constant and the electron filling in the *p* orbital perpendicular to the molecular plane. This method is applied to 100 molecules containing a ^{14}N quadrupolar nucleus using data from more than 40 years of rotational spectroscopy and comparing the performance of three levels of theory from quantum-chemical computations. From this extensive dataset we have analyzed chemical properties of molecules, such as the hybridization of the nucleus, and we could quantify the extent of polarization and resonance processes as well as physical characteristics of the quadrupolar nucleus, such as *eQq*. This is a constant, which represents the coupling in the hypothetical case of having a single electron in an electronic orbital at the isolated nucleus, and its value has been under long debate. Here, the *eQq* value has been determined for the ^{14}N nucleus, and the methodology to calculate it for other nuclei is provided.

I. INTRODUCTION

"Nuclear quadrupole coupling constants in molecules depend on the nuclear quadrupole moments and the variation in electrostatic field at the nucleus". With this quote starts the canonical paper from Townes and Dailey about nuclear quadrupole effects and their relationship with the electronic structure of molecules.[1] The Townes-Dailey (TD) model postulates that nuclear electric quadrupole interaction is caused by the coupling of the electric field gradient (EFG) from electrons in the valence shell *p* orbitals (vide infra) with the nuclear quadrupole moment (Q) of atoms with a nuclear spin I>1/2.[1] Therefore the quadrupole interaction has an electronic origin, and the nuclear quadrupole coupling constants (NQCCs) in molecules can be related to the electronic structure of the molecule, chemical bonds, orbital hybridization of the quadrupolar nucleus or resonance processes.[2,3-4]

In order to obtain reliable information about the electronic structure or the chemical bonding, it is critical to understand how the NQCCs scale with the electronic charge. Actually, NQCCs are proportional to the unpaired number of electrons in the atomic orbitals related by a parameter (χ_0 or eQq_{n10}, vide infra) that englobes the elementary electric charge, *Q,* and the EFG (q_{n10}).[1-3] An eQq_{n10} has a constant value for each quadrupolar nucleus, since it represents the coupling due to a single unpaired electron in a *p* orbital for the isolated nucleus. In $eQq_{n10,}$ the elementary charge *e* is well-known,[5] the nuclear quadrupole moment *Q* is specific for each nucleus and its values are tabulated,[3] and q_{n10} is the EFG, which can be computed from the molecular wave functions. In principle, it seems as all the components are known and deduction of electronic molecular properties from NQCCs is a rudimentary labor. In reality, though, this is a more challenging task; computing the EFG is complicated since it requires considering the contribution of all the charges, and only an approximation to

the wave functions can be computed. Approximations used to speed up calculations, like pseudopotentials, sacrifice accuracy in predicting the EFG. This is particularly true for late periodic table elements like iodine and gold.[6,7] As a result, some eQq_{n10} values seem to be not reliable, or different values have been reported with large discrepancies between them. For example, for ^{127}I, eQq_{510} was tabulated as 292.71 MHz,[3] but it was mistyped, with the actual value being 2292.71 MHz.[8] In the case of ^{14}N there is no agreement on its value to date. First estimations of eQq_{210} (^{14}N) ranged from -10 to -24 MHz.[1] This value was later limited to be between -10 and -11 MHz.[9-10,11] A more modern update by quantum-beat and radio-frequency resonance spectroscopy reported it to be |11.2| MHz.[12] Most of contemporary authors take the negative sign from the first data, with the updated value of 11.2 MHz.

In this work, we have focused on ^{14}N-containing molecules to explore the relationship between molecular structure, NQCCs, and eQq_{n10}, aiming at information about molecular properties and at deriving an accurate, reliable eQq_{210} value for the ^{14}N atom from experimental data.

II. THEORETICAL BACKGROUND

In the TD model, NQCCs (χ_{ij}, *i,j* = *x, y, z*) arise from the coupling of the nuclear quadrupole moment (Q) for those nuclei with a non-spherical distribution of the charge, i.e., with a nuclear spin I >1/2, with the EFG (q_{ij}) at the nucleus ($\chi_{ij} = eQq_{ij}$). To estimate q_{ij} it would be needed, in principle, to consider the EFG due to the charges in the molecule over the coupling nucleus (q_{mol}), but the TD model considers some approximations to simplify the calculation.[2,3] EFG scales with $1/r^3$, so only the electrons of the coupling nucleus have a significant influence ($q_{mol} = q_{nlml}$). The electrons in the core filled shells and those in the valence shell *s* orbitals are spherically symmetric so their

contribution can be neglected. The contribution of d electrons has been estimated to be lower than those in a p orbital by a factor of 47 and is thus omitted here.[8] In the space-fixed axis system, the maximum effective value for the electric field gradient occurs for the quantum numbers $l = 1$ and $m_l = 0$. Thus, the electric field gradient that produces nuclear quadrupole coupling can be approximately attributed to an uneven filling of the valence shell p orbitals of the coupling atom ($q_{nlml} = q_{n10}$).[2,3] With these approximations, the EFG at the coupling atom caused by a single electron is:

$$q_x = -\left(\frac{n_y + n_z}{2} - n_x\right)q_{n10} \qquad \text{(Eq. 1)}$$

$$q_y = -\left(\frac{n_z + n_x}{2} - n_y\right)q_{n10} \qquad \text{(Eq. 2)}$$

$$q_z = -\left(\frac{n_x + n_y}{2} - n_z\right)q_{n10} \qquad \text{(Eq. 3)}$$

The term inside brackets represents the unequal filling of the p orbitals with n_i ($i = x$, y or z) being the number of electrons in the p_i orbital, which can be computed for example from Natural Bond Orbital calculations (NBO).[13,14] This term is more frequently represented as $-(U_p)_i$ ($i = x$, y or z) and represents the electron density deficit or excess along the i axis at the coupling nucleus. If the population of the p_z orbital is greater than the semi sum of the population of p_x and p_y, ($n_z > \left(\frac{n_x + n_y}{2}\right)$) then $-(U_p)_z$ will be positive, indicating an excess of electron density along the z axis. On the other hand, if $n_z < \left(\frac{n_x + n_y}{2}\right)$, $-(U_p)_z$ will be negative, indicating a deficit of electron density along the z axis.

The original TD model has been extended[8] and modified[15] to improve its performance. Recently, Rinald and Wu have introduced the concept of valence p-orbital population anisotropy (VPPA) in their modified Townes-Dailey model, showing how the unequal filling of p-orbitals is related to the NQC tensor and the χ_0 constant. Their VPPA term is equivalent to our $-(U_p)_z$ since both illustrate the unbalanced (or unpaired) number of electrons in a p orbital, based on the electronic populations from natural molecular orbital calculations. We will continue using $-(U_p)_z$ and eQq_{210} considering that it is the classical nomenclature defined by Townes and Dailey and most of the researchers in rotational spectroscopy may be familiar with this terminology.

Substituting q_z from equation 3 with $\chi_{ij} = eQq_{ij}$ gives:

$$\chi_{zz} = eQ\left(-\left(\frac{n_x + n_y}{2} - n_z\right)q_{n10}\right) = eQ - \left(U_p\right)_z q_{n10}$$

(Eq. 4)

$$\frac{\chi_{zz}}{eQq_{n10}} = -\left(U_p\right)_z \qquad \text{(Eq. 5)}$$

From those equations it is evident how sensitive the NQCC values are to the electronic environment, making them unique for each molecule, isomer and even conformers.[2,3,16,17] Experimentally, the measurable NQCCs are defined in the molecular inertial axis system, $\chi_{\alpha\beta}$ ($\alpha,\beta = a$, b or c), but eQq_{n10} is directly related to the NQCCs in the

quadrupolar atom axis system (χ_{ij}). Transformation between both axis systems will require the diagonalization of the quadrupole coupling tensor, which can only be performed if the tensor is completely known. For certain nuclei, which are strongly coupled and thus show rather large quadrupole coupling splitting, such as some halogens, it is usually possible to determine the whole tensor experimentally. For the weaker coupling nuclei, such as ^{14}N, the complete NQC tensor can often not be determined, since the effect of the off-diagonal elements is almost negligible or the intensity of the transitions are below the detection limits of the instrumentation. Missing terms can be taken from quantum-chemical calculations to complete the experimental tensor, or in another approximation, the theoretical tensor can be diagonalized and then the transformation matrix employed is applied to the experimental values to obtain approximated χ_{ij} values. Nevertheless, for certain molecular systems this process may not be needed. Provided the planarity or almost planarity of a molecular system, the perpendicular NQCC will be parallel to one of the χ_{ii} constants, and from eq. 5 it can be correlated with the corresponding $-(U_p)_i$. The association between $\chi_{\alpha\beta}$ and χ_{ij} is not always straightforward since there are several ways of correlating the molecular axis system with the quadrupolar nuclei axis system.[3] The NBO calculations performed in this work indicate that for all the molecules studied the $2p_z$ orbital is perpendicular to the molecular plane. Based on that, in this work we have considered that the NQCC perpendicular to the molecular plane, χ_{cc}, is approximately equivalent to χ_{zz}. In this case, it can be correlated with $-(U_p)_z$, which represents the unbalanced electronic charge from the perpendicular $2p_z$ orbital (Figure 1). This approximation has been used in other studies,[18,19-20] where relevant information was concluded from the evolution of χ_{cc} for a series of complexes, and is extended to a variety of molecules here.

In rotational spectroscopy, the nuclear quadrupole coupling interaction is observed as a hyperfine structure of the rotational transitions in the spectra. Rotational spectrometers are characterized by a high resolution, which involves the measurement of accurate frequencies for the rotational and hyperfine transitions. This is translated into an accurate determination of the NQCCs. Thus, rotational spectroscopy is an excellent technique to look for data from which to obtain an accurate value for the eQq_{n10} parameter. The aim of this work is to provide proof of principle of this approach and to apply it to an extensive data set of the available molecules in the bibliography. This method could be extended later to more molecular systems and other quadrupolar nuclei.

III. METHODOLOGY

A total of 108 data points for ^{14}N NQCCs have been used for this study. They have been obtained from a bibliographical search of a plethora of studies covering more than 40 years of rotational spectroscopy from many research groups devoted to Fourier transform microwave (FTMW) spectroscopy. Different experimental set-ups based on this technique were used; Fabry-Perot cavity FTMW spectrometers,[21] or broadband FTMW instruments.[22] The bibliographical search for experimental data points was guided by the collection of molecules at Prof. Bailey's website.[23] Table S1 in the supplementary information collects the experimental rotational and quadrupole coupling constants together with P_{cc} and χ_{cc}/eQq_{210}. The approximation employed ($\chi_{cc} \approx \chi_{zz} = -(U_p)_z \cdot eQq_{n10}$) is only

valid for planar systems. The planarity of a molecular or complex can be evaluated by the P_{cc} parameter, which gives the mass extension out of the molecular plane; the greater P_{cc}, the less planar the system. We have observed that even for complexes with a planar moment of up to 16 uÅ^2, such as the complexes with three molecules with either ethyl carbamate or phenanthridine (Table S1), the approximation is still valid. The complete list of references for the experimental values is also given in the supplementary information.
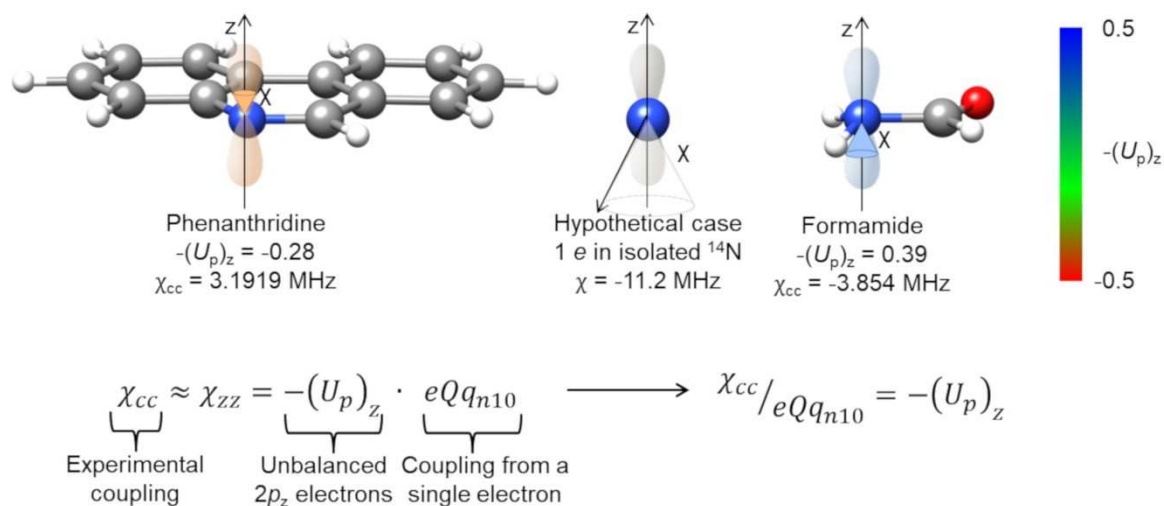

Phenanthridine
$-(U_p)_z = -0.28$
$\chi_{cc} = 3.1919$ MHz

Hypothetical case
1 e in isolated ^{14}N
$\chi = -11.2$ MHz

Formamide
$-(U_p)_z = 0.39$
$\chi_{cc} = -3.854$ MHz

$$\underbrace{\chi_{cc}}_{\substack{\text{Experimental} \\ \text{coupling}}} \approx \chi_{zz} = \underbrace{-(U_p)_z}_{\substack{\text{Unbalanced} \\ 2p_z \text{ electrons}}} \cdot \underbrace{eQq_{n10}}_{\substack{\text{Coupling from a} \\ \text{single electron}}} \longrightarrow \chi_{cc}/eQq_{n10} = -(U_p)_z$$

Figure 1 The approximation employed in this work and the correlation between χ_{cc}, $-(U_p)_z$, and eQq_{n10}. Representation of the quadrupole coupling in phenanthridine and formamide based on their $-(U_p)_z$ and the hypothetical case of an electron in an isolated ^{14}N nucleus.
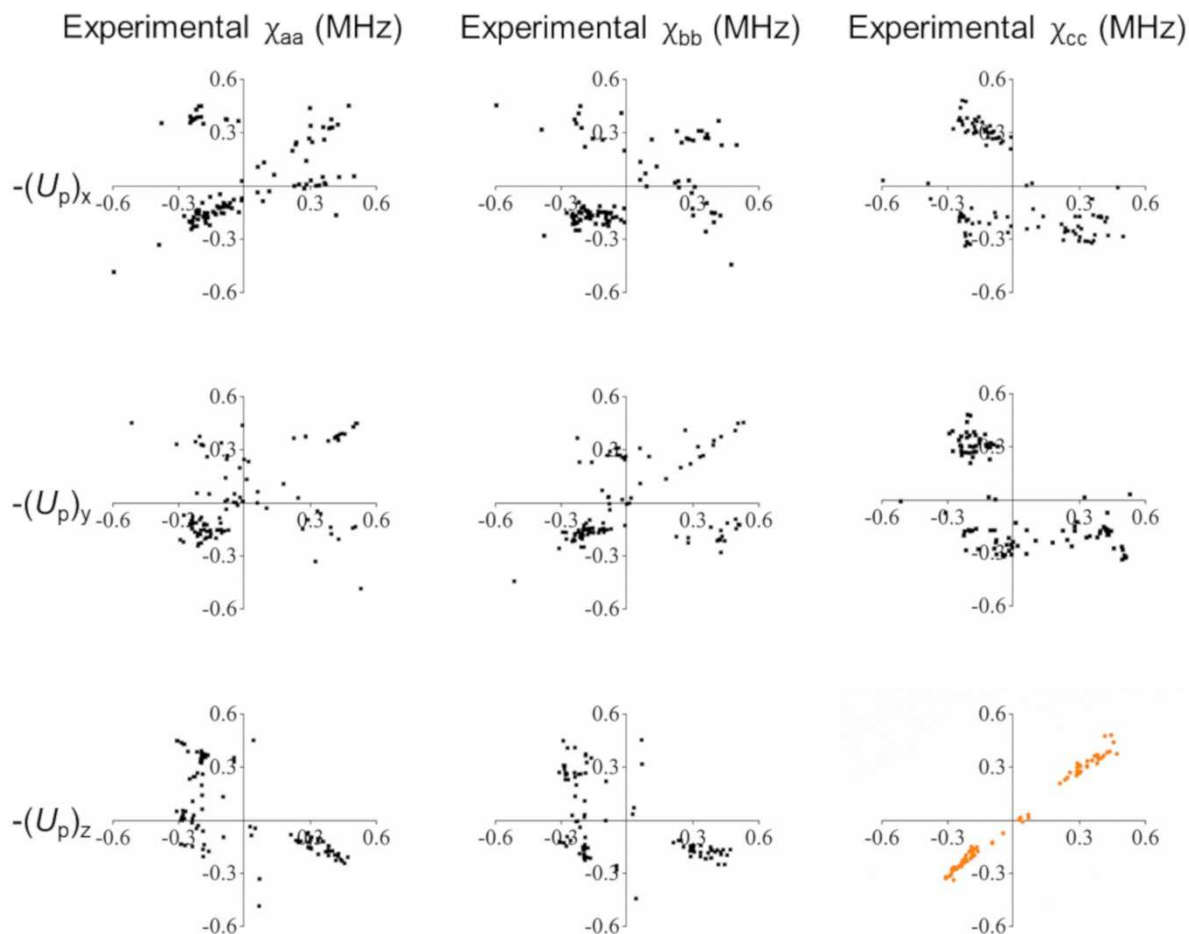
Figure 2. The nine possible combinations of experimental $\chi_{\alpha\alpha}$ ($\alpha = a$, b or c) versus -$(U_p)_i$ ($i = x$, y or z) in matrix arrangement. All the experimental quadrupole values are divided by eQq_{210} taken as -11.2 MHz.

The calculations to obtain $–(U_p)_z$ were performed at three levels of theory: B3LYP[24]/cc-PVDZ[25], B3PW91[14b,26]/6-311+G(df,pd),[27] and MP2[28]/cc-PVDZ. B3LYP has been chosen based on its popularity and relatively low computational cost,[29] B3PW91/6-311+G(df,pd) has been reported to be a specially good level of theory to describe ^{14}N quadrupolar momentum,[30] and MP2 was chosen to provide a non-DFT value to compare. In some sections of this study, only the results at the B3PW91/6-311+G(df,pd) level of theory are given since it is the recommended method,[30] but the results from the three levels of theory are comparable. The first step was to perform structure optimizations using the Gaussian suite;[31] after that we performed NBO calculations using the optimized geometry to obtain the electron filling of the valence-shell orbitals. The accuracy of the quantum-chemical calculations was tested by comparison of the predicted rotational constants and NQCCs with the experimental data given in the bibliography. For every level of theory there are some data points that do not reproduce the experimental information or for which the calculation did not converge into a stable structure. Those data points were excluded from the analysis for their tendency line (vide infra). Tables S2-S4 collect the predicted parameters for the three levels of theory employed indicating which data points were excluded.

IV. RESULTS AND DISCUSSION

To corroborate the validity of equation 6, we compared the experimental $\chi_{\alpha\alpha}/eQq_{210}$ ($\alpha = a$, b or c) versus the calculated -$(U_p)_i$ ($i = x$, y or z) at the three levels of theory for 108 data points of [14]N. At this stage eQq_{210} was taken as -11.2 MHz, [8,12] which is a good starting value despite the inconsistencies of previous studies. There are nine possible -$(U_p)_i$ versus $\chi_{\alpha\alpha}$ combinations, which have been represented in the matrix arrangement for the B3PW91/6-311+G(df,pd) level of theory (Figure 2). The strong correlation between -$(U_p)_z$ and χ_{cc} confirms that the approximation $\chi_{cc} \approx \chi_{zz} = -(U_p)_z \cdot eQq_{210}$ is valid for the molecular systems chosen. A linear feature but with higher dispersion of the data seems to appear for the pairs, -$(U_p)_x/\chi_{aa}$ and -$(U_p)_y/\chi_{bb}$, although it could be expected to observe a higher correlation degree after the diagonalization of the quadrupole tensor.

In Figure 3, the experimental χ_{cc}/eQq_{210} is compared with the calculated -$(U_p)_z$ at the three levels of theory, showing that the three quantum-chemical methods give comparable results without a significant difference between them. Some systematic effects can be observed, i. e., the theoretical NQCCs values are predicted to be larger than the experimental ones for the included aniline-like molecules (aniline, methylaniline, dimethylaniline, 4-aminobenzonitrile and 4,4'-dimethylaminobenzonitrile, Tables S2-S4).

The values do not cover the complete line, but are concentrated between the absolute values (0.45 – 0.2) with just a few examples out of this range and close to 0. Since the values are related with the electron filling of the nitrogen $2s$ and $2p$ orbitals, a possible explanation could be that only certain values of the electron filling might give rise to stable molecules. The molecules and complexes studied cover the different hybridization possibilities for the nitrogen atom; sp, sp^2, and sp^3, and they can be classified according to their position (Figure 4). The negative quadrant (negative value for χ_{cc}/eQq_{210}; positive χ_{cc}) contains sp molecules and those with sp^2 hybridization, in which the lone pair of nitrogen is localized in one of the sp^2 orbitals and the remaining non-hybridized $2p$ orbital is involved in a resonance process. Examples of molecules in this quadrant are pyridine, oxazole, benzonitrile, and molecules with similar chemical environments. The molecules with values close to 0 are 3-methylisoxazole, 4-methylisoxazole, 5-methylisoxazole, isocyanobenzene, and pentafluoroisocyanobenzene. For these molecules, the nitrogen atom has a similar hybridization as the molecules in the negative quadrant, but it is attached to a highly electronegative group. The positive quadrant (positive χ_{cc}/eQq_{210}; negative χ_{cc}) includes molecules with sp^3 and sp^2 hybridization, in which the lone pair from the nitrogen atom is localized in the non-hybridized $2p$ orbital. Data points in this section belong to amines, amides, and pyrrole- or indole-like molecules. There are also examples of molecules with two nitrogen atoms with different hybridization schemes, such us imidazole or 4-aminobenzonitrile. Another, simpler classification is possible. Molecules in the negative quadrant or close to 0 belong to data points in which the nitrogen is not attached to a hydrogen atom, while molecules in the positive quadrant have an N-H bond.
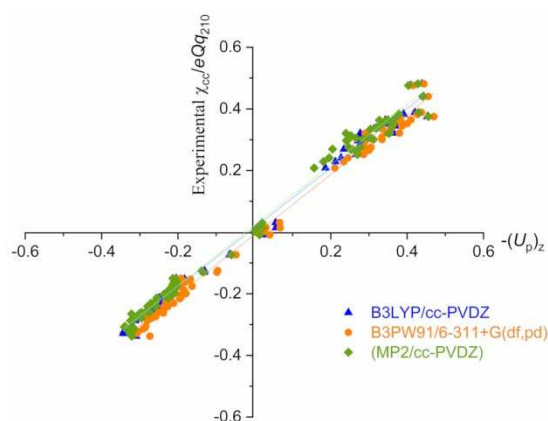

Figure 4. Experimental χ_{cc}/eQq_{210} versus calculated $-(U_p)_z$ at the B3PW91/6-311+G(df,pd) level of theory illustrating the different hybridization for the nitrogen atom in the different quadrants. eQq_{210} was taken as -11.2 MHz.

Polarization and resonance effects in complexes

From the definition of $-(U_p)_z$ given above, a negative value as in phenanthridine (PAN) corresponds to an electron density deficit at the ^{14}N nucleus along the z axis, while a positive value as in formamide indicates an excess of electron density.[3] For both cases, the monomer and complexes with up to three molecules of water have been measured,[18-20,32] and the values of χ_{cc} show a steady tendency (either an increase or a decrease) with the hydration degree (Figure 5). The change in χ_{cc} can be associated with the electronic environment at the ^{14}N nucleus and therefore with the polarization via resonance processes. In the series of microsolvated complexes of PAN, χ_{cc}/eQq_{210} increases with the number of water molecules, which corresponds to an increase in the electron density along the z axis at the nitrogen nucleus.[20] In the series of microsolvated complexes of formamide, χ_{cc} increases as more molecules of water are added, and consequently, as a result of the negative sign of eQq_{210}; χ_{cc}/eQq_{210} decreases. That illustrates the decrease in the electron density excess with cluster size (Figure 5).[18,19] The tendency of χ_{cc} for both systems is in good agreement with the polarization expected from the resonance forms that can be deduced from the electronic delocalization. In PAN, the resonance of the π-system creates a partial negative charge at the nitrogen nucleus (Figure S1). The resonance can be extended to the water molecules, and therefore as the cluster increases its size, there is more resonance and the electron density at the nitrogen atom increases. In formamide, resonance induces a partial positive charge at the nitrogen atom (Figure S1). As the cluster size increases, the electron density at the nitrogen atom decreases. The experimental values of χ_{cc} reproduce perfectly the evolution of the electron density on both molecular systems (Figure 5). Similar tendencies in χ_{cc} have been observed also for monomers and complexes with one water molecule of pyridine, benzonitrile, indole, pyrrole, azetidinone, and ethyl carbamate (Figure S2). Although the data set is more limited, the change in χ_{cc} agrees with the expected behavior according to their resonance forms.

Figure 3. Experimental χ_{cc}/eQq_{210} versus calculated $-(U_p)_z$ at the B3LYP/cc-PVDZ, B3PW91/6-311+G(df,pd) and MP2/cc-PVDZ levels of theory. eQq_{210} was taken as -11.2 MHz.
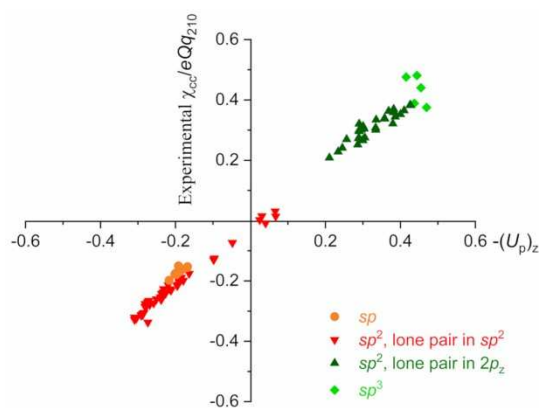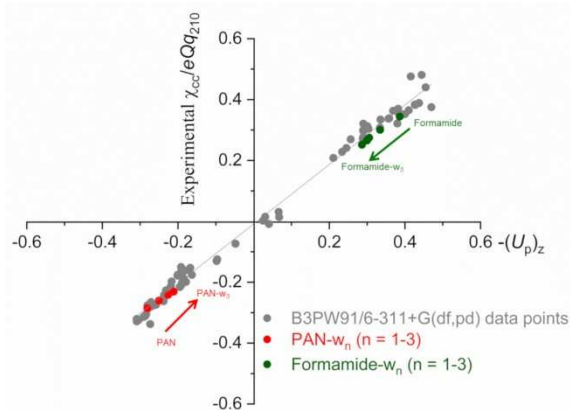
Figure 5. Experimental χ_{cc}/eQq_{210} versus $-(U_p)_z$ at B3PW91/6-311+G(df,pd). The data points corresponding to complexes with up to three molecules of water for phenanthridine and formamide have been highlighted in red and green, respectively. The changes in χ_{cc}/eQq_{210} and $-(U_p)_z$ from the monomer to complexes with increasing water molecules are represented by the arrows. eQq_{210} was taken as -11.2 MHz.

Electron density and resonance effects in aromatic rings

Resonance can be evidenced by trends in the NQCCs not only in series of complexes, but also in molecules with different substituents. In aromatic rings, the influence of substituents on the electron density of the ring is well known. The substituents can have two effects, inductive or mesomeric, and both can be either positive (electron donating) or negative (electron withdrawing). Depending on the atom and its relative position, the substituent affects the electron density of the respective atoms in the ring differently and in turn affects the molecular structure. This is an important principle for synthetic chemistry, and the substituents and their effects are well characterized. The quantification of the effects of electron-donating or -withdrawing groups on molecular structures is more challenging. Several studies have reported the observation of structural changes by means of microwave spectroscopy in aromatic rings.[33-34,35,36,37,38] Experimentally, the determination of molecular structures relies on the observation of every possible mono-isotopic substitution, either in natural abundance or using enriched samples, which can be a challenge. The trends in the values of the NQCC χ_{cc} of ^{14}N in heteroaromatic rings can provide further confirmation of resonance and allow us to trace back the electron density even if the molecular structures are not determined.

In the following, the approach will be exemplified for pyridine (pyr) and fluoro-substituted pyridines (F-pyr).[35,36] Fluorine is a common substituent of aromatic rings, and it induces inductive and mesomeric effects, but with different sign. It extracts electronic charge from the ring since it is more electronegative than the carbon atoms in the aromatic ring. In addition, it donates electronic charge through resonance of the lone pairs (Figure S3 for resonance structures). The consequence on the electronic charge results from the addition of both effects and depends on the position of the fluorine substituent in the ring. The relationship between the experimental χ_{cc}/eQq_{210} and the calculated $-(U_p)_z$ at the B3PW91/6-311+G(df,pd) level of theory of pyr and the F-pyr molecules follows the described tendency line (Figure 6 and Figure S4 for a zoom-in). Fluorine substitution in different positions with respect to the nitrogen atom affects the electron density at the ^{14}N atom differently, resulting in changes of the χ_{cc} values. The experimental values of the NQCC can be used as a probe to quantify the amount of electron density delocalization in molecules with one or also multiple substituents, for which the effect adds up.
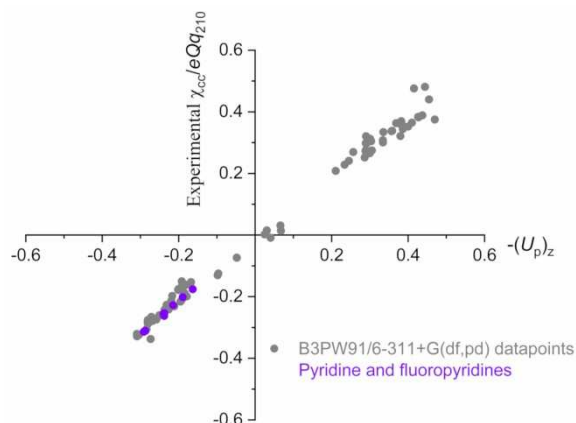

Figure 6. Experimental χ_{cc}/eQq_{210} versus $-(U_p)_z$ at B3PW91/6-311+G(df,pd). The data points corresponding to pyr and F-pyr have been highlighted in purple. eQq_{210} was taken as -11.2 MHz.

Table 1 shows experimental χ_{cc}, the calculated $2p_z$ electron filling, and $-(U_p)_z$ computed at the B3PW91/6-311+G(df,pd) level of theory. Table S6 collects the values at the three levels of theory. The effect of fluorine substitution in the three possible positions has been calculated from the mono-substituted molecules, and then those values have been used to extrapolate the same properties for difluoro- and pentafluoropydirine. The extrapolated values for experimental χ_{cc}, $2p_z$ electron filling, and $-(U_p)_z$ seem to match with the experimental ones. The net effect of fluorine in the 2-position (ortho-) will be to donate 0.04 electrons to the $2p_z$ orbital, affecting the χ_{cc} value by -0.638 MHz. This change is preserved in every system with a fluorine atom in the 2-position, even if more fluorine atoms are present. In 4- (para-) and 3- (meta-) positions, fluorine will donate 0.02 or extract around 0.007 electrons to the $2p_z$ orbital, respectively. By direct comparison of the experimental χ_{cc}, we can trace back the electron density and illustrate the electron donating or withdrawing processes. The change in the electronic population of the $2p_z$ orbital could be in principle derived from the Extended Townes-Dailey (ETD) analysis,[8,39] which also requires some approximations and assumptions based on quantum-chemical calculations. The application of the ETD is out of the scope of this work but similar conclusions could be expected.

Table 1. Experimental χ_{cc}, predicted and extrapolated $2p_z$ electron filling, and $-(U_p)_z$ at B3PW91/6-311+G(df,pd) for pyridine, monofluoropyridines, difluoropyridines, and pentafluoropyridine.

	Experimental χ_{cc} (MHz)	Theoretical (B3PW91)		Fluorine effect			Extrapolated fluorine effect		
		$2p_z$	$-(U_p)_z$	$\Delta\chi_{cc}$ (MHz)	$\Delta 2p_z$	$\Delta-(U_p)_z$	χ_{cc} (MHz)	$2p_z$	$-(U_p)_z$
Pyridine	3.474	1.160	-0.288	-	-	-	-	-	-

2-Fluoropyridine	2.837	1.203	-0.238	-0.638[b]	0.043	0.050	-	-	-
3-Fluoropyridine	3.501	1.153	-0.290	0.027	-0.007	-0.002	-	-	-
4-Fluoropyridine[a]	2.997	1.180	-0.265	-0.477	0.020	0.023	-	-	-
2,3-Difluoropyridine	2.940	1.199	-0.238	-	-	-	2.863[c]	1.196	-0.240
2,4-Difluoropyridine	2.548	1.222	-0.215	-	-	-	2.360	1.223	-0.215
2,5-Difluoropyridine	2.851	1.197	-0.240	-	-	-	2.863	1.196	-0.240
2,6-Difluoropyridine	2.259	1.244	-0.190	-	-	-	2.199	1.246	-0.188
3,4-Difluoropyridine[a]	3.025	1.173	-0.267	-	-	-	3.024	1.173	-0.267
3,5-Difluoropyridine	3.519	1.148	-0.291	-	-	-	3.527	1.146	-0.292
Pentafluoropyridine	1.970	1.257	-0.163	-	-	-	1.775	1.252	-0.169

[a] No experimental value available, only theoretical. [b] Contribution from fluorine in the 2-, 3-, or 4- position to χ_{cc}, $2p_z$ or $-(U_p)_z$. [c] Value calculated from that of non-substituted pyridine and the corresponding contributions from monofluoropyridines.

eQq

Besides the validation of the method and obtaining information about electron density through the experimental NQCC, one important aim of this work is a derivation of a semi-experimental value for the eQq_{210} parameter for ^{14}N. In previous sections we have taken eQq_{210} = -11.2 MHz,[8] but we can also determine this critical parameter from the selected dataset. If we represent directly the experimental values of χ_{cc} instead of χ_{cc}/eQq_{210} versus $-(U_p)_z$, , the gradient of the line should be eQq_{210} (Figure 7). To obtain the value of the gradient we performed a linear fit of the data points constraining the line to pass through 0. The experimental regression line does not pass exactly through the 0.0 point (Figure 7), but in the fit we set that constrain assuming that no unbalanced $2p_z$ electrons would correspond to 0 χ_{cc}. We performed the fit for the calculated values of $-(U_p)_z$ at the three levels of theory and compared the results.

The eQq_{210} values obtained are -10.79(10), -10.93(10), and -10.90(13) MHz with r^2 of 0.991, 0.991 and 0.984 using B3LYP/cc-PVDZ, B3PW91/6-311+G(df,pd), and MP2/cc-PVDZ, respectively. The average from the three levels of theory will be eQq_{210} = -10.87(11) MHz, which is within the range given previously (-10 to -11 MHz), but with higher accuracy in its determination and with the negative sign as suggested by most of contemporary authors.[3,8] To try to reduce the uncertainty we have also performed the fit separating the dataset in two groups; monomers and complexes, since the last could, in principle, present a higher deviation based on their higher P_{cc} values (Figure S5). The result of this fit is comparable to the one including every molecular system, indicating that although complexation may affect the planarity of the system, the approximation $\chi_{cc} \approx \chi_{zz} = -(U_p)_z \cdot eQq_{210}$ is still valid.
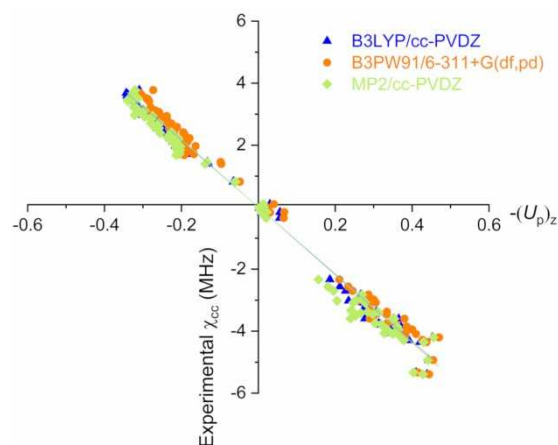

Figure 7. Experimental χ_{cc} versus calculated $-(U_p)_z$ at the B3LYP/cc-PVDZ, B3PW91/6-311+G(df,pd) and MP2/cc-PVDZ levels of theory. The gradient of the regression line corresponds to eQq_{210}.

V. CONCLUSIONS

In this work, we provide a methodology to extract information about nuclear and molecular properties from the experimental NQCC χ_{cc} values together with some straightforward quantum-chemical calculations, with the only requirement that the molecular system is planar or near-planar. χ_{cc} can be determined with high accuracy in a straightforward manner from high-resolution FTMW spectroscopy. With the methodology provided in the first section of this work, it is possible to stablish a relationship between χ_{cc} and the unbalanced p_z electronic charge. This methodology was applied to a plethora of molecular systems containing ^{14}N. The values and tendencies in the NQCC χ_{cc} evidenced polarization processes in a series of complexes and allowed to trace out the electron density in substituted aromatic rings. The most outstanding conclusion from this work is the derivation of an accurate semi-experimental value for eQq_{210} for ^{14}N from a simple linear fit. The exact value of this parameter was still an open question for physical chemistry. Our extensive dataset and the good agreement of the three levels of theory result in a value of eQq_{210} = -10.87(11) MHz. The same methodology, based on comparison of experimental χ_{cc} values obtained from a bibliographical search and simple quantum-chemical calculations using low-cost levels of theory can be easily extended to other quadrupolar nuclei to obtain their correspondent eQq_{n10} values.

SUPPLEMENTARY MATERIAL

The supplementary material includes experimental and calculated parameters at the three levels of theory values for the molecular systems included in this work. Resonance structures and additional plots are also included. The complete list of references from which values for this work have been taken is collected in the Supporting Information.

ACKNOWLEDGMENTS

DATA AVAILABILITY

The data used for this study can be found in this article, in the supplementary material and from the authors upon reasonable request.

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